

13 NOV 1968
MDDC-887

13 NOV 1968
MDDC-887

UNITED STATES ATOMIC ENERGY COMMISSION

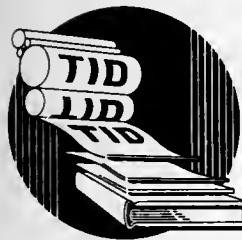
MDDC-887

ISOTOPIC MASSES AND ABUNDANCES (LADC-285)

By
H. A. Bethe
R. F. Christy

July 21, 1943

Los Alamos Scientific Laboratory



Technical Information Division, ORE, Oak Ridge, Tennessee



Date Declassified: March 6, 1947

Issuance of this document does not constitute authority for declassification of classified copies of the same or similar content and title and by the same authors.

Reproduced direct from copy
as submitted to this office.

PRINTED IN USA
PRICE 5 CENTS

ISOTOPIC MASSES AND ABUNDANCES*

By H. A. Bethe and R. F. Christy

This table contains the isotopic mass values which we consider most reliable. As can be seen from the given probable errors, the masses up to neon are very much more accurately known than for the heavier elements. The lighter masses were calculated, taking into account all of the more accurate data from mass spectrograph as well as disintegrations, and attempting the best possible fit.

Abbreviations used in the isotopic mass table:

For stable isotopes, the abundance is given in per cent. Radioactive elements are indicated by the particle they emit ($\beta-$ or $\beta+$); βc means that the nucleus captures an orbital electron but does not emit positrons. P means instability against disintegration into heavy nuclear particles. Mass values, in parenthesis, refer to nuclei which have not yet been produced.

The errors listed are meant to be about 3 times the probable or twice the standard error. For theoretical estimates, the error is usually not given.

Sources of the data are indicated as follows:

I—Calculations by Betty J. Isaacs, Cornell University, Master's Thesis, 1942.

W—Wigner, Memorandum No. 24, April 2, 1940.

B—Barkas, Phys. Rev. 55:691 (1939).

M—Mass spectrograph value (usually most reliable).

D—From disintegration experiments involving heavy particles from end-point of β -spectrum.

C—Calculated value.

corr—corrected in view of more recent accurate determination of the mass of a neighboring isotope, from which the mass of the given isotope is obtained through a disintegration measurement or a theoretical estimate.

E—Estimate.

* This is Section 19 of LA 11, Los Alamos Handbook of Nuclear Physics.

Z	Element	A	Abundance per cent	Mass	Error x 10 ⁵	Source
0	N	1	β^-	1.00893	3	ID
1	H	1	99.93	1.00813	0.6	IM
1	H	2	.02	2.014708	1.1	IM
1	H	3	β^-	3.01702	3.4	ID
2	He	3	10^{-5}	3.01700	4	ID
2	He	4	100	4.00390	3	IM and D
2	He	5	P	5.0137	35	ID
2	He	6	β^-	6.0209	50	IB
3	Li	5	P	(5.0136)	(60)	WC
3	Li	6	7.9	6.01697	5	ID
3	Li	7	22.1	7.01822	6	IM and D
3	Li	8	β^-	8.02502	7	D
4	Be	6	β^+	6.0119	(100)	WC
4	Be	7	-c	7.01916	7	ID
4	Be	8	P	8.00785	7	ID
4	Be	9	100	9.01503	6	ID
4	Be	10	β^-	10.01677	8	ID
4	Be	11	P	(11.0277)	-	WC
5	B	9	P	9.01620	7	ID
5	B	10	18.4	10.01618	9	IM and D
5	B	11	81.6	11.01284	8	IM and D
5	B	12	β^-	12.0190	70	ID
5	B	13	β^-	(13.0207)	-	WC
6	C	10	β^+	10.0210	30	ID
6	C	11	β^+	11.01495	9	ID
6	C	12	98.9	12.00382	4	IM
6	C	13	1.1	13.00751	10	IM and D
6	C	14	β^-	14.00767	5	ID and β
6	C	15	β^-	(15.0165)	-	WC
7	N	12	(β^+)	(12.0233)	-	WC
7	N	13	β^+	13.00988	7	ID
7	N	14	99.62	14.00751	4	IM

Z	Element	A	Abundance per cent	Mass	Error $\times 10^5$	Source
7	N	15	0.38	15.00489	21	IM
7	N	16	β^-	$\left\{ \begin{array}{l} > 16.0065 \\ < 16.011 \end{array} \right\}$	-	β ID
7	N	17	β^-	(17.014)	-	WC
8	O	14	β^+	(14.0131)	-	WC
8	O	15	β^+	15.0078	40	$I\beta$
8	O	16	99.76	16.000000	-	Standard
8	O	17	0.04	17.00450	6	ID
8	O	18	0.20	18.0049	40	WM
8	O	19	(β^-)	19.0139)	-	WC
9	F	16	β^+	(16.0175)	-	WC
9	F	17	β^+	17.0075	30	ID
9	F	18	β^+	18.0065	60	WD
9	F	19	100	19.00450	26	IM
9	F	20	β^-	$\begin{array}{l} > 20.0042 \\ < 20.0092 \end{array}$	-	$I\beta$ ID
9	F	21	β^-	(21.0059)	-	WC
10	Ne	18	β^+	(18.0114)	-	WC
10	Ne	19	β^+	19.00781	20	$W\beta$
10	Ne	20	90.0	19.99877	10	IM
10	Ne	21	0.27	20.99963	22	IM
10	Ne	22	9.73	21.99844	36	IM
10	Ne	23	β^-	(23.0013)	-	WC
11	Na	21	β^+	(21.0035)	-	WC
11	Na	22	β^+	21.9999	50	$I\beta$
11	Na	23	100	22.99618	31	ID
11	Na	24	β^-	23.9975	45	ID
11	Na	25	β^-	(24.9967)	-	WC
12	Mg	22	β^+	(22.0062)	-	WC
12	Mg	23	β^+	23.0002	40	$W\beta$
12	Mg	24	77.4	23.9924	60	$B\beta$
12	Mg	25	11.5	24.9938	90	WD
12	Mg	26	11.1	25.9898	50	WD

Z	Element	A	Abundance per cent	Mass	Error $\times 10^5$	Source
12	Mg	27	β^-	26.9928	150	W β
13	Al	25	β^+	24.9981	100	W β
13	Al	26	β^+	25.9929	150	W β
13	Al	27	100	26.9899	80	WD
13	Al	28	β^-	27.9903	70	W β
13	Al	29	β^-	28.9893	80	W β
13	Al	30	β^-	(29.9954)	-	WC
14	Si	27	β^+	26.9949	90	I β
14	Si	28	89.6	27.9866	60	WM
14	Si	29	6.2	28.9866	60	WM
14	Si	30	4.2	29.9832	90	WD
14	Si	31	β^-	30.9862	60	W β
14	Si	32	β^-	(31.9849)	-	WC
15	P	29	β^+	(28.9919)	(100)	WC
15	P	30	β^+	29.9873	100	I β
15	P	31	100	30.9843	50	WM
15	P	32	β^-	31.9827	40	I β
15	P	33	β^-	(32.9826)	-	WC
16	S	31	β^+	(30.9899)	-	C
16	S	32	95.0	31.98089	7	IM
16	S	33	0.74	32.9800	60	W corr
16	S	34	4.2	33.97710	35	IM
16	S	35	β^-	34.9788	80	W corr
16	S	36	0.016	35.978	100	W
17	Cl	33	β^+	(32.9860)	-	WC corr
17	Cl	34	β^+	33.9801	-	I β
17	Cl	35	75.4	34.97867	21	IM
17	Cl	36	βc	35.9788	100	W
17	Cl	37	24.6	36.97750	14	IM
17	Cl	38	β^-	37.981	300	W
17	Cl	39	β^-	(38.9794)	-	WC
18	A	35	β^+	(34.9850)	-	C

Z	Element	A	Abundance per cent	Mass	Error x 10 ⁵	Source
18	A	36	0.307	35.9780	100	W
18	A	37	β^-	(36.9777)	-	E
18	A	38	0.061	37.974	250	ID
18	A	39	β^-	(38.9755)	-	WC
18	A	40	99.63	39.9756	60	IM
18	A	41	β^-	40.9770	60	W
19	K	37	β^+	(36.9830)	-	WC
19	K	38	β^+	(37.9795)	-	WC
19	K	39	93.3	(38.9747)	-	WC
19	K	40	0.012	39.9760	100	W
20	Ca	40	96.96	39.9753	150	E
20	Ca	42	0.64	41.9711	-	W
20	Ca	43	0.15	42.9723	-	W
20	Ca	44	2.06	-	-	-
20	Ca	45	β^-	44.968	-	G
21	Sc	45	100	44.9669	-	ID
22	Ti	46	-	45.9661	100	IM
22	Ti	47	-	46.9647	100	IM
22	Ti	48	-	47.9631	50	IM
22	Ti	49	-	48.9646	60	IM
22	Ti	50	-	49.9621	40	IM
22	Ti	51	-	50.9587	100	W corr
23	V	51	-	50.9577	50	-
24	Cr	51	β^-	50.958	-	W corr
24	Cr	52	81.6	51.956	-	W corr
24	Cr	53	10.4	52.956	-	W corr
25	Mn	55	100	54.957	-	E
26	Fe	54	6.04	53.957	-	W corr
26	Fe	56	91.57	55.9568	170	IM
26	Fe	57	2.11	56.957	-	W corr
26	Fe	58	0.28	-	-	-

UNIVERSITY OF FLORIDA



3 1262 08909 7991